

## Useful Formulas

$\text{kg H}_2\text{SO}_4 = 0.98 \times \text{kg CaCO}_3$  - Note that in Australia, acid base accounting calculations are based on the net acidity of samples ( $\text{kg of H}_2\text{SO}_4/\text{t}$ ), whereas in North America it is based on the net neutralizing potential available ( $\text{kg of CaCO}_3/\text{t}$ ).

$\text{pyrite}\% = \text{sulfur}\% \times 120/64$

$\text{sulfur}\% = \text{pyrite}\% \times 64/120$

$\text{carbon}\% \times 81.66 = \text{kg H}_2\text{SO}_4/\text{t}$  neutralizing capacity (assuming that all the carbon is calcium carbonate)

$\text{MPA (kg H}_2\text{SO}_4/\text{t}) = 30.59 \times \text{sulfur}\%$  (assuming the sulfide is pyrite)

$\text{ANC (kg H}_2\text{SO}_4/\text{t}) = 0.98 \times \text{kg CaCO}_3/\text{t}$

1 ounce = 28.35 gram

1 kilogram = 2.2 pound

1 tonne = 1.1 ton

1 metre = 3.28 feet

1 kilometre = 0.62 mile

1 hectare = 2.47 acres

1 litre = 0.264 gallon

1 cubic metre = 35.3 cubic feet